**General Chem: Review for Test 4**

The test will consist of the following:

Six of these vocabulary words:

* alkali metal
* alkaline earth metal
* transition metal
* actinide
* lanthanide
* halogen
* noble gas
* malleable
* ductile
* cation
* anion
* polyatomic ion
* ion
* octet rule
* crystal lattice
* unit cell

Three of these questions:

* How are ionic bonds formed?
* Why are ionic compounds hard?
* Why are ionic compounds brittle?
* Why do ionic compounds have high melting and boiling points?
* Why do ionic compounds conduct electricity when melted or dissolved?
* Why don’t ionic compounds burn?

Eight of each of the following type of question (16 total):

* You should know how to name ionic compounds if given the formulas. You will be given access to the two-sided periodic table form on the test
  + NaOH, Pb(OH)2, SnO2, CaF2, (NH4)2CO3, Na2SO4 and other similar compounds.
* You should know how to write ionic formulas if given the names. You will be given access to the two-sided periodic table form on the test
  + lithium acetate, iron (III) phosphate, potassium phosphide, vanadium (III) sulfite, and similar compounds.